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| Periodic Table Project  Evaluators Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | | | | | | | | |
| Group Members: | |  |  |  |  |  |  |  |
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| Periodic Table Group Name | | | NOBLE GASES | |  |  |  |  |
|  |  |  |  |  |  |  |  |  |
| Number of Valence Electrons (6pts) | | |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |
| Element Information (18pts)  One point per section | | Metal / Nonmetal / Metalloid | | Phase | Fact |  |  |  |
|  | Helium |  | |  |  |  |  |  |
|  | Neon |  | |  |  |  |  |  |
|  | Argon |  | |  |  |  |  |  |
|  | Krypton |  | |  |  |  |  |  |
|  | Xenon |  | |  |  |  |  |  |
|  | Radon |  | |  |  |  |  |  |
|  |  |  |  |  |  |  | Sub Total |  |
| Trends(36pts)  3pts per section | |  | Definition | | Trend |  |  |  |
|  | Atomic Mass | |  | |  |  |  |  |
|  | Atom Size | |  | |  |  |  |  |
|  | Reactivity | |  | |  |  |  |  |
|  | Shielding Effect | |  | |  |  |  |  |
|  | Ionization Energy | |  | |  |  |  |  |
|  | Electronegativity | |  | |  |  |  |  |
|  |  |  |  |  |  |  | Sub Total |  |
|  |  |  |  |  |  |  |  |  |
| Sources (10pts) |  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |
| Creativity & Neatness (10pts) |  |  |  |  |  |  |  |  |
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| Board/Power Point Display (20pts) | | | |  |  |  |  |  |
|  | All Previous information present | | | | 5pts |  |  |  |
|  | Use of Academic Language | | |  | 5pts |  |  |  |
|  | Draw the attention of the viewer | | | | 5pts |  |  |  |
|  | Answers Questions about this group to understand their trends | | | | | | | 5pts |
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| Comments: | |  |  |  |  |  | **TOTAL** |  |

Extra Credit Points (5pts each)

Did student answer the following questions:

Why are they called Nobel Gases?

Answer should be (or similar to) : Because Valence electrons are full

Why is Helium considered a Noble Gas:

Answer should be (or Similar to): Because Helium has only one Valence shell and because the first Valence shell can only hold two electrons, it’s valence shell is full therefore it is a Noble Gas.

Questions you must answer to receive credit for this exam

**Only need to answer it in your self-evaluation form, but they must be answered or you only receive a 50 for your test grade.**

1. What trends can you find when evaluating your and the other posters/presentation

Atomic Mass:

Atomic Size/Radius:

Reactivity:

Shielding effects:

Ionization Energy:

Electronegativity:

1. What do all the elements in your group have in common: Name two things

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| Periodic Table Project | | | | | | | | |
| Group Members | |  |  |  |  |  |  |  |
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| Periodic Table Group Name | | | HALOGENS | |  |  |  |  |
|  |  |  |  |  |  |  |  |  |
| Number of Valence Electrons (6pts) | | |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |
| Element Information (18pts) | | Metal / Nonmetal / Metalloid | | Phase | Fact |  |  |  |
|  | Fluorine |  | |  |  |  |  |  |
|  | Chlorine |  | |  |  |  |  |  |
|  | Bromine |  | |  |  |  |  |  |
|  | Iodine |  | |  |  |  |  |  |
|  | Astatine |  | |  |  |  |  |  |
|  |  |  |  |  |  |  | Sub Total |  |
| Trends(36pts) | |  | Definition | | Trend |  |  |  |
|  | Atomic Mass | |  | |  |  |  |  |
|  | Atom Size | |  | |  |  |  |  |
|  | Reactivity | |  | |  |  |  |  |
|  | Shielding Effect | |  | |  |  |  |  |
|  | Ionization Energy | |  | |  |  |  |  |
|  | Electronegativity | |  | |  |  |  |  |
|  |  |  |  |  |  |  | Sub Total |  |
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| Sources (10pts) |  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |
| Creativity & Neatness (10pts) |  |  |  |  |  |  |  |  |
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| Board/Power Point Display (20pts) | | | |  |  |  |  |  |
|  | All Previous information present | | | | 5pts |  |  |  |
|  | Use of Academic Language | | |  | 5pts |  |  |  |
|  | Draw the attention of the viewer | | | | 5pts |  |  |  |
|  | Answers Questions about this group to understand their trends | | | | | | | 5pts |
|  |  |  |  |  |  |  |  |  |
| Comments: | |  |  |  |  |  | **TOTAL** |  |

Extra Credit Points (10 pts each)

Did student answer the following questions:

What happens to electrons when these elements form ionic compounds?

Answer should be (or similar to): They have a negative charge or become anions or they pick up an electron, so that their valence shell can have 8 electrons allowing them to have properties more like that of Noble Gases.

Questions you must answer to receive credit for this exam

**Only need to answer it in your self-evaluation form, but they must be answered or you only receive a 50 for your test grade.**

1. What trends can you find when evaluating your and the other posters/presentation

Atomic Mass:

Atomic Size/Radius:

Reactivity:

Shielding effects:

Ionization Energy:

Electronegativity:

What do all the elements in your group have in common: Name two things

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| Periodic Table Project | | | | | | | | |
| Group Members | |  |  |  |  |  |  |  |
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| Periodic Table Group Name | | | ALKALINE EARTH METALS | | | |  |  |
|  |  |  |  |  |  |  |  |  |
| Number of Valence Electrons (6pts) | | |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |
| Element Information (18pts) | | Metal / Nonmetal / Metalloid | | Phase | Fact |  |  |  |
|  | Beryllium |  | |  |  |  |  |  |
|  | Magnesium |  | |  |  |  |  |  |
|  | Calcium |  | |  |  |  |  |  |
|  | Strontium |  | |  |  |  |  |  |
|  | Barium |  | |  |  |  |  |  |
|  | Radium |  | |  |  |  |  |  |
|  |  |  |  |  |  |  | Sub Total |  |
| Trends(36pts) | |  | Definition | | Trend |  |  |  |
|  | Atomic Mass | |  | |  |  |  |  |
|  | Atom Size | |  | |  |  |  |  |
|  | Reactivity | |  | |  |  |  |  |
|  | Shielding Effect | |  | |  |  |  |  |
|  | Ionization Energy | |  | |  |  |  |  |
|  | Electronegativity | |  | |  |  |  |  |
|  |  |  |  |  |  |  | Sub Total |  |
|  |  |  |  |  |  |  |  |  |
| Sources (10pts) |  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |
| Creativity & Neatness (10pts) |  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |
| Board/Power Point Display (20pts) | | | |  |  |  |  |  |
|  | All Previous information present | | |  | 5pts |  |  |  |
|  | Use of Academic Language | | |  | 5pts |  |  |  |
|  | Draw the attention of the viewer | | |  | 5pts |  |  |  |
|  | Answers Questions about this group to understand their trends | | | | | |  | 5pts |
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| Comments: | |  |  |  |  |  | **TOTAL** |  |
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Extra Credit Points (10 pts each)

Did student answer the following questions:

What happens to electrons when these elements form compounds?

Answer should be (or similar to): They have a positive charge of 2+ or become cations or they readily loose two valence electrons to form a positive ion or cation .

Questions you must answer to receive credit for this exam

**Only need to answer it in your self-evaluation form, but they must be answered or you only receive a 50 for your test grade.**

1. What trends can you find when evaluating your and the other posters/presentation

Atomic Mass:

Atomic Size/Radius:

Reactivity:

Shielding effects:

Ionization Energy:

Electronegativity:

What do all the elements in your group have in common: Name two things

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| Periodic Table Project | | | | | | | | |
| Group Members | |  |  |  |  |  |  |  |
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| Periodic Table Group Name | | | ALKALI METALS | | |  |  |  |
|  |  |  |  |  |  |  |  |  |
| Number of Valence Electrons (6pts) | | |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |
| Element Information (18pts) | | Metal / Nonmetal / Metalloid | | Phase | Fact |  |  |  |
|  | Lithium |  | |  |  |  |  |  |
|  | Sodium |  | |  |  |  |  |  |
|  | Potassium |  | |  |  |  |  |  |
|  | Rubidium |  | |  |  |  |  |  |
|  | Cesium |  | |  |  |  |  |  |
|  | Francium |  | |  |  |  |  |  |
|  |  |  |  |  |  |  | Sub Total |  |
| Trends(36pts) | |  | Definition | | Trend |  |  |  |
|  | Atomic Mass | |  | |  |  |  |  |
|  | Atom Size | |  | |  |  |  |  |
|  | Reactivity | |  | |  |  |  |  |
|  | Shielding Effect | |  | |  |  |  |  |
|  | Ionization Energy | |  | |  |  |  |  |
|  | Electronegativity | |  | |  |  |  |  |
|  |  |  |  |  |  |  | Sub Total |  |
|  |  |  |  |  |  |  |  |  |
| Sources (10pts) |  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |
| Creativity & Neatness (10pts) |  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |  |
| Board/Power Point Display (20pts) | | | |  |  |  |  |  |
|  | All Previous information present | | | | 5pts |  |  |  |
|  | Use of Academic Language | | | | 5pts |  |  |  |
|  | Draw the attention of the viewer | | | | 5pts |  |  |  |
|  | Answers Questions about this group to understand their trends | | | | | | | 5pts |
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| Comments: | |  |  |  |  |  | **TOTAL** |  |
|  |  |  |  |  |  |  |  |  |

Extra Credit Points (10 pts each)

Did student answer the following questions:

Why is Hydrogen in this column?

Answer should be (or similar to): Because Hydrogen even though it is not an Alkaline Metal, it is in this column because it only has one valence electron.

What happens to electrons when these elements form compounds?

Answer should be (or similar to): They tend to donate their electrons in reactions and have an oxidation state of +1

Questions you must answer to receive credit for this exam

**Only need to answer it in your self-evaluation form, but they must be answered or you only receive a 50 for your test grade.**

1. What trends can you find when evaluating your and the other posters/presentation

Atomic Mass:

Atomic Size/Radius:

Reactivity:

Shielding effects:

Ionization Energy:

Electronegativity:

What do all the elements in your group have in common: Name two things