Name**\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_(SHOW YOUR WORK)**

1. How many grams of O2 is produced if 2.50 g of KClO3 is completely decomposed by heating?

2 KClO3 (s) 🡪 2 KCl (s) + 3 O2 (g)

2.5 g x g

x g O2 = 2.5 ~~g KClO~~~~3~~ x (1 mol KClO3) x (3 ~~mol O~~~~2~~) x (32 g O2) = 0.98 g O2

(122.5 ~~g KClO~~~~3~~) (2 mol KClO3) (1 ~~mol O~~~~2~~)

2. How many grams of KCl is produced if 2.50 g of KClO3 is decomposed?

1. How many moles of KClO3 is used to produce 10 moles of O2?
2. How many moles of KCl is produced if 15 g of KClO3 is used?
3. How many liters of O2 is produced if 5 moles of KClO3 is used?
4. How many liters of O2 is produced if 10 g of KClO3 s used?